

(19)



Europäisches Patentamt

European Patent Office

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(11)

EP 0 919 695 A2

(12)

EUROPEAN PATENT APPLICATION

(43) Date of publication:
02.06.1999 Bulletin 1999/22

(51) Int. Cl.⁶: E21B 43/26

(21) Application number: 98122163.3

(22) Date of filing: 25.11.1998

(84) Designated Contracting States:
AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE
Designated Extension States:
AL LT LV MK RO SI

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(30) Priority: 25.11.1997 CA 2222087

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(54) Encapsulated breaker for oil gel system

(57) An encapsulated breaker for a fracturing fluid for use in fracturing subterranean formations comprising: a hydrocarbon base; neutralized alkyl phosphate

esters complexed with metallic cations, to form a gel, in said hydrocarbon base.

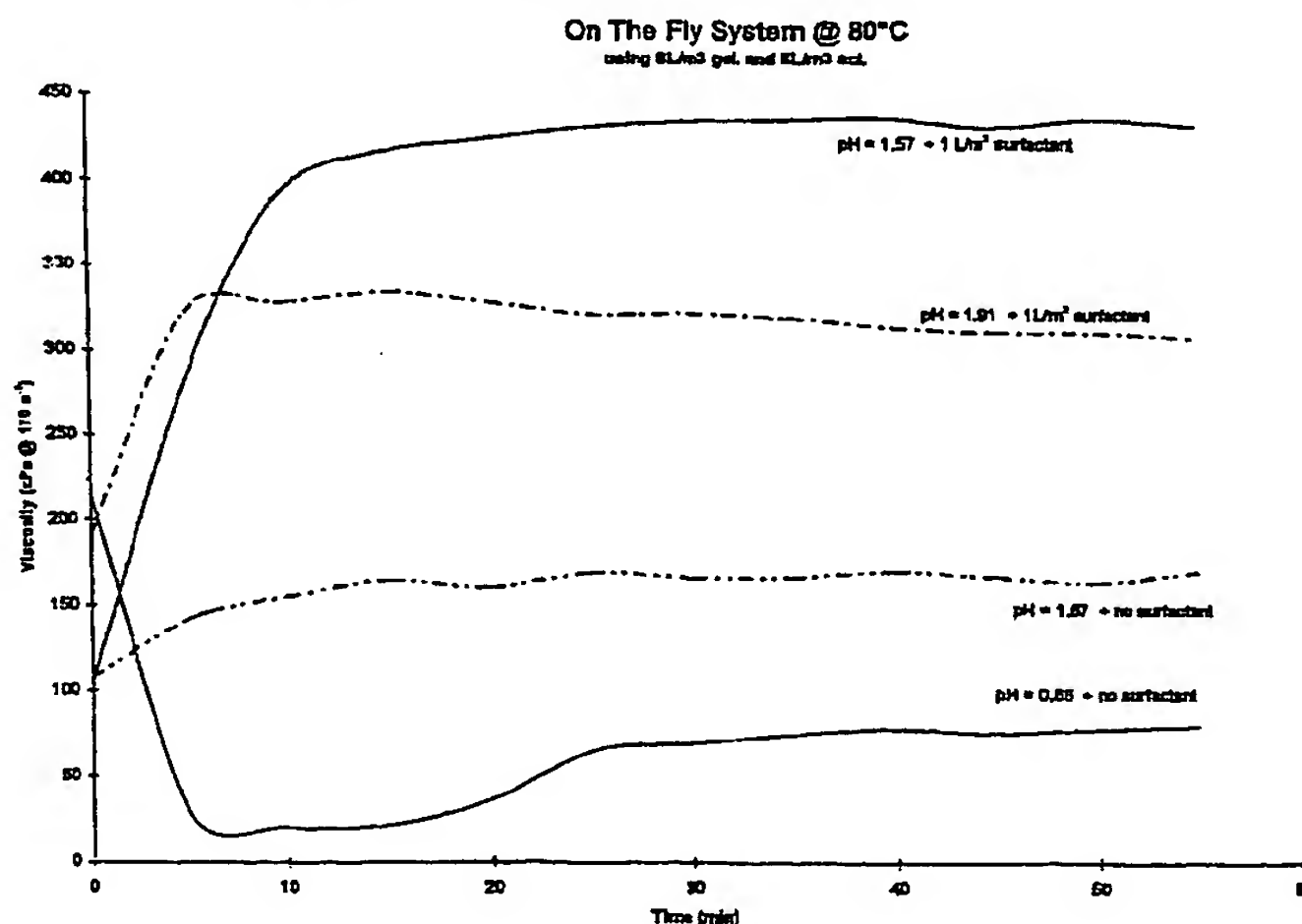


FIGURE 1

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Description

[0001] The present invention relates to methods and compositions for fracturing subterranean formations. In particular, the present invention describes a novel encapsulated breaker for use with a fracturing fluid such as those described in copending Canadian patent application No. 2237383.

[0002] In order to increase the productivity of oil and gas wells, or to bring back into production wells that have essentially gone dry, it is common practice to conduct a procedure known as fracturing the well. In such a procedure, fluids known as fracturing fluids, are injected into the well at very high hydrostatic pressures. The fluids are typically viscous gels, and act under pressure to open pores and cracks in the formation, thereby to increase the overall permeability of the well. Typically, the fluids also are used to transport proppants, such as sand or ground walnut shells into the cracks and pores, thereby to ensure that the pores and cracks formed during fracturing remain open under the lower pressures that will exist after fracturing, when the well is producing.

[0003] The fluid that has been used to fracture the formation is removed by the introduction of viscosity lowering agents into same, which permit the fluid to be more easily pumped from the formation. These agents are known as breakers because they tend to break down the fracturing gel. Breakers can act on a gel in a number of ways, such as by the random oxidation of polymers to shorten the chain length thereof. In the present invention, a breaker is utilized to adjust the pH of a gel, to break same by hydrolysis.

[0004] The fracturing fluid of the present invention is a phosphate alkyl ester gel. It is known that a gel can be produced by mixing trivalent cations, such as aluminum, with a phosphate alkyl ester. However, gels obtained with known phosphate alkyl esters have not been commercially acceptable, because the viscosity developed with same has been insufficient or slow to develop, and usually required tank mixing.

[0005] Phosphate alkyl esters may be mono-esters, di-esters or tri-esters. In the mono-ester, one primary mono-hydric alcohol, of C₅-C₁₆ length is ester linked to a phosphate. A di-ester has two such alkyl alcohols linked to a phosphate. A tri-ester has three alkyl alcohols linked to it. The applicants have discovered that commercially valuable gels are feasible with a di-ester content above 50%, preferably above 65%, and a tri-ester content below 5%. The remainder may be mono-ester.

[0006] The applicants enhance gel development in two important ways. First, the phosphal alkyl esters are neutralized with primary amines. Secondly the applicants utilize a surfactant to enhance gel development. An appropriate surfactant is ammonium alkyl (C₆-C₂₀) sulfate.

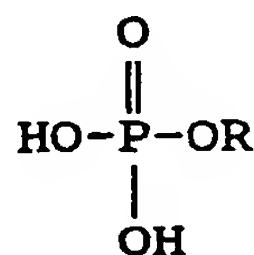
[0007] In order to break the gel described herein the preferred breaker of the present invention is a pH adjusting breaker, such as calcium oxide and sodium carbonate. It is preferred that the breaker be encapsulated in a porous inert substance, such as nylon.

[0008] In a broad aspect, then, the present invention relates to an encapsulated breaker for a fracturing fluid for use in fracturing subterranean formations comprising: a hydrocarbon base; and neutralized alkyl phosphate esters complexed with metallic cations, to form a gel, in said hydrocarbon base.

[0009] In drawings that illustrate the present invention by way of example:

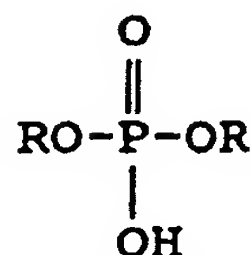
Figure 1 is a graph illustrating the effect of pH and surfactant on the phosphate alkyl ester gels of the present invention.

[0010] The phosphate ester gels of the present invention are made from primary mono-hydric alcohols of C₅-C₁₆ chain length. It is preferred to utilize alcohols of chain length C₈-C₁₂, but it will be understood that since the gels of the present invention are intended to act on hydrocarbon fluids, such as diesel fuel, kerosene, or other common hydrocarbon fluids, the selection of an optimal chain length alcohol will be a matter of choice for one skilled in the art. The alkyl alcohols are combined with phosphates, by known techniques to produce mono-di and tri-alkyl esters which may generally be represented as follows:



mono-alkyl phosphate ester

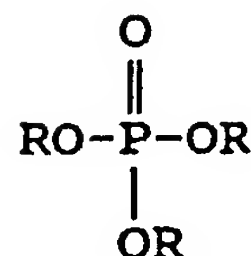
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di-alkyl phosphate ester

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tri-alkyl phosphate ester

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[0011] The relative proportions of mono-alkyl, di-alkyl and tri-alkyl esters are important. The applicant has discovered that the di-alkyl phosphate ester content must be at least 50% for any gel development to occur, and must be at least 65% for commercially feasible gel to form. Moreover, tri-alkyl phosphate ester content must be lower than 5%, or poor gel rheological characteristics will result. There must be at least a small percentage of mono-alkyl phosphate ester present to initiate gelling.

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[0012] Primary amines are used to neutralize esters. The quantity of amine used will, of course depend on the amine. However, a suitable amine is mono-ethanol amine (MEA). From 1% to 5% of a concentrated (99%) MEA solution (relative to the weight of the phosphate alkyl esters) is used to produce a desired level of neutralization (pH) in the esters. A preferred pH is about 1.4-1.8.

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[0013] The partially neutralized phosphate ester is then mixed in the hydrocarbon fluid to be gelled, with a trivalent cation, either aluminum or ferric. It is, moreover, preferred to utilize chelated ions, since these readily disperse in organic solvents, thereby enhancing the rate of gelation. Examples of suitable compounds are aluminum isopropoxide, aluminum sec-butoxide, oxy-aluminum octoate, and so on.

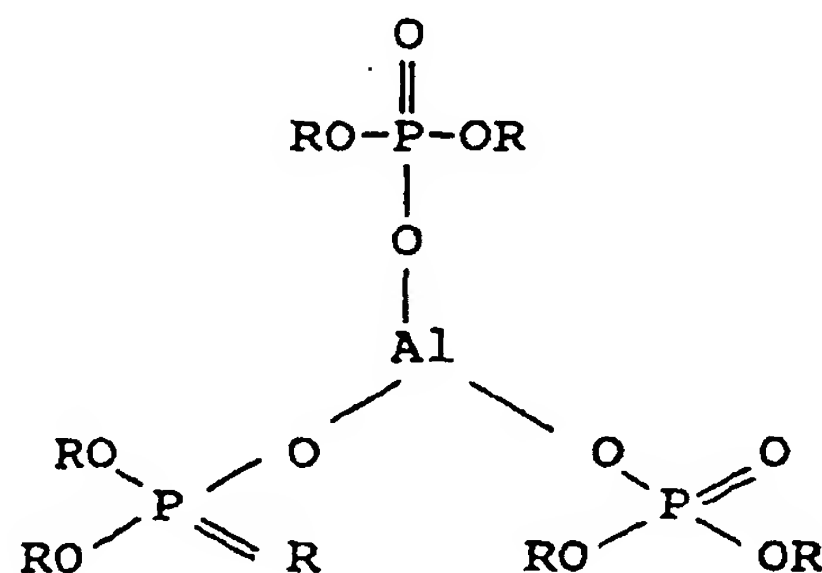
[0014] The phosphate-alkyl esters and cations form a mesh-like network in the hydrocarbon fluid, resulting in a gel. The hypothesized reaction is as follows:

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[0015] It will be understood that such a network will transform the essentially two dimensional phosphate-alkyl ester to a three dimensional, very viscous gel.

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[0016] The gelling rate can moreover be greatly enhanced using a surfactant. Sulphates or sulphonate groups provide the necessary alteration of surface chemistry to result in enhancement of reaction rates. Aluminum alkyl (C₆-C₂₀) sulfate has been found to be an appropriate surfactant.

[0017] The effects of partial neutralization and the use of surfactants are illustrated in Figure 1. In the examples shown in Figure 1, it can be seen that the lowest viscosity is achieved with an unneutralized gel, and no surfactant. It will be understood that in each case 8 litres per cubic metre diesel oil, of a 70% dialkyl phosphate ester are utilized, with

8l/cubic metre of aluminum tri-isopropoxide.

[0018] In the example neutralized to a pH of 1.57, resulting viscosity is about doubled, with the use of a surfactant resulting in a further tripling of the achieved viscosity. As shown, if the pH is permitted to rise, viscosity will be lowered somewhat. This information can be utilized, however, to advantage, to design gels with viscosity characteristics to particularly match those desired for fracturing a given formation.

[0019] To form a gel in a hydrocarbon solvent according to the present invention, 0.5 - 1.5% (by weight of hydrocarbon) partially neutralized phosphate alkyl ester are utilized with 0.025 - 0.2% surfactant and 0.1 - 1.5% trivalent cation. About 80% of the maximum viscosity of the gel is attained, typically, in the first minute of mixing. The constituents can and advantageously are, therefore, blended "on the fly" as they are pumped into a formation. Aluminum acetate may also be added, to stabilize the gel.

[0020] The gel may be broken by the use of pH adjusting breakers, such as soda ash, caustic, lime, amines, and acids.

[0021] Particular preferred breakers are calcium oxide, and calcium peroxide, encapsulated in a particle of about 80-100 mesh size. The breaker is encapsulated by known techniques, such as a fluidized bed process, or interfacial polymerization. The thickness, and permeability, of the membrane of the breaker can be adjusted, by known techniques, to provide any desired release rate of breaker.

[0022] The effect of encapsulated calcium oxide breaker, on a gel described herein, is summarized in Table I which follows. The procedure used to generate the data of Table I is as follows:

1. The gellant (OG-14X Gellant) at 0.8 vol.% loading is added to the hydrocarbon fluid (diesel) while stirring at 1200 rpm with a lab stirrer.
2. Activator (OG-14X Activator) at 0.8 vol.% loading is added to the fluid.
3. Mixing speed is increased to 1500 rpm.
4. After 15 sec. mixing (to mimic suction manifold to tub time), the stabilizer (A5) at 0.025 wt% is added followed by the encapsulated breaker of this invention.
5. Mixing speed is increased to 2500 rpm and the fluid is mixed for additional 45 sec.
6. The fluid is placed in the high temp. rheometer (Fann 50) and tested.

[0023] The procedure was then repeated, with various concentrations of water present, to simulate various well conditions.

TABLE I

96-331XO RESULTS

SAMPLE	A			B			C		
	μ	n^1	K^1	μ	n^1	K^1	μ	n^1	K^1
INITIAL	278	0.25	27.12	240	0.17	28.17	200	0.20	23.72
AFTER SCAN	330	--	--	245	--	--	270	--	--
10 min	435	0.17	51.95	405	0.08	61.72	185	0.07	31.23
20 min	440	0.15	54.58	395	0.06	62.69	150	0.09	22.17
30 min	440	0.13	57.76	370	0.06	60.80	105	0.18	11.74
max visc.	445	--	--	410	--	--	370	--	--

SAMPLE	D			E			F		
	μ	n^1	K^1	μ	n^1	K^1	μ	n^1	K^1
INITIAL	164	0.14	21.55	98	0.96	1.09	169	0.33	13.61
AFTER SCAN	189	--	--	120	--	--	186	--	--
10 min	200	0.02	33.3	410	0.35	28.09	390	0.22	38.69
20 min	135	0.11	19.58	435	0.13	58.71	405	0.11	57.15
30 min	95	0.23	9.77	425	0.14	52.73	410	0.10	57.36
max visc.	280	--	--	455	--	--	410	--	--

SAMPLE	G			H			I		
	μ	n'	K'	μ	n'	K'	μ	n'	K'
INITIAL	190	0.22	20.78	273	0.24	27.72	191	0.27	--
AFTER SCAN	205	--	--	378	--	--	197	--	--
10 min	370	0.10	53.83	255	0.12	33.15	370	0.19	40.5
20 min	410	0.10	59.48	120	0.22	12.08	320	0.24	32.14
30 min	400	0.10	59.33	55	0.44	2.79	280	0.30	23.27
max visc.	410	--	--	490	--	--	370	--	--

J		
SAMPLE	μ	K'
INITIAL	200	0.72
AFTER SCAN	380	--
10 min	185	0.60
20 min	245	0.49
30 min	275	0.42
max visc.	430	--

DESCRIPTION

TEMP(°C)

SAMPLE

A	No breaker	80
B	96-331XO @ 2.0 (No H ₂ O)	60
C	96-331XO @ 2.0 (1% H ₂ O)	60
D	96-331XO @ 2.0 (5% H ₂ O)	60
E	94-44X @ 3.0 OG Breaker B @ 0.3 *NOIHCBI	80
F	96-331XO @ 4.0 (0.5% H ₂ O)	35
G	96-331XO @ 4.0 (0.5% H ₂ O)	35
H	96-331XO @ 2.0 (1.0 H ₂ O)	80
I	96-331XO @ 2.0 (No H ₂ O)	80
J	OG Breaker B @ 2.0 (No H ₂ O)	60

P-40 OG-14X Gel @ 8 OG-14X Act @ 8 IHCBI @ 0.25 (±)

n' is a unitless constant (power law coefficient)
 k' is expressed as (Pa.s) ^{n'} for a power law fluid.

[0024] It will be appreciated, therefore, that the present invention provides a reliable viscous gel and encapsulated breaker for use with hydrocarbon solvents, which can be broken on a consistent and effective basis.

Claims

1. An encapsulated breaker for a fracturing fluid for use in fracturing subterranean formations comprising:

- 5 (a) a hydrocarbon base;
 (b) neutralized alkyl phosphate esters complexed with
 (c) metallic cations, to form a gel, in said hydrocarbon base.

2. The fracturing fluid as claimed in claim 1, wherein said encapsulated breaker is a peroxide or oxide.

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3. The fracturing fluid as claimed in claim 2, wherein the breaker is encapsulated calcium oxide.

4. The fracturing fluid as claimed in claim 1, 2 or 3, wherein said breaker is encapsulated in a porous, inert encapsulant.

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5. The fracturing fluid as claimed in claim 4, wherein said encapsulant is nylon.

6. The fracturing fluid as claimed in claim 5, wherein said encapsulated breaker has an average size of 80 - 100 mesh.

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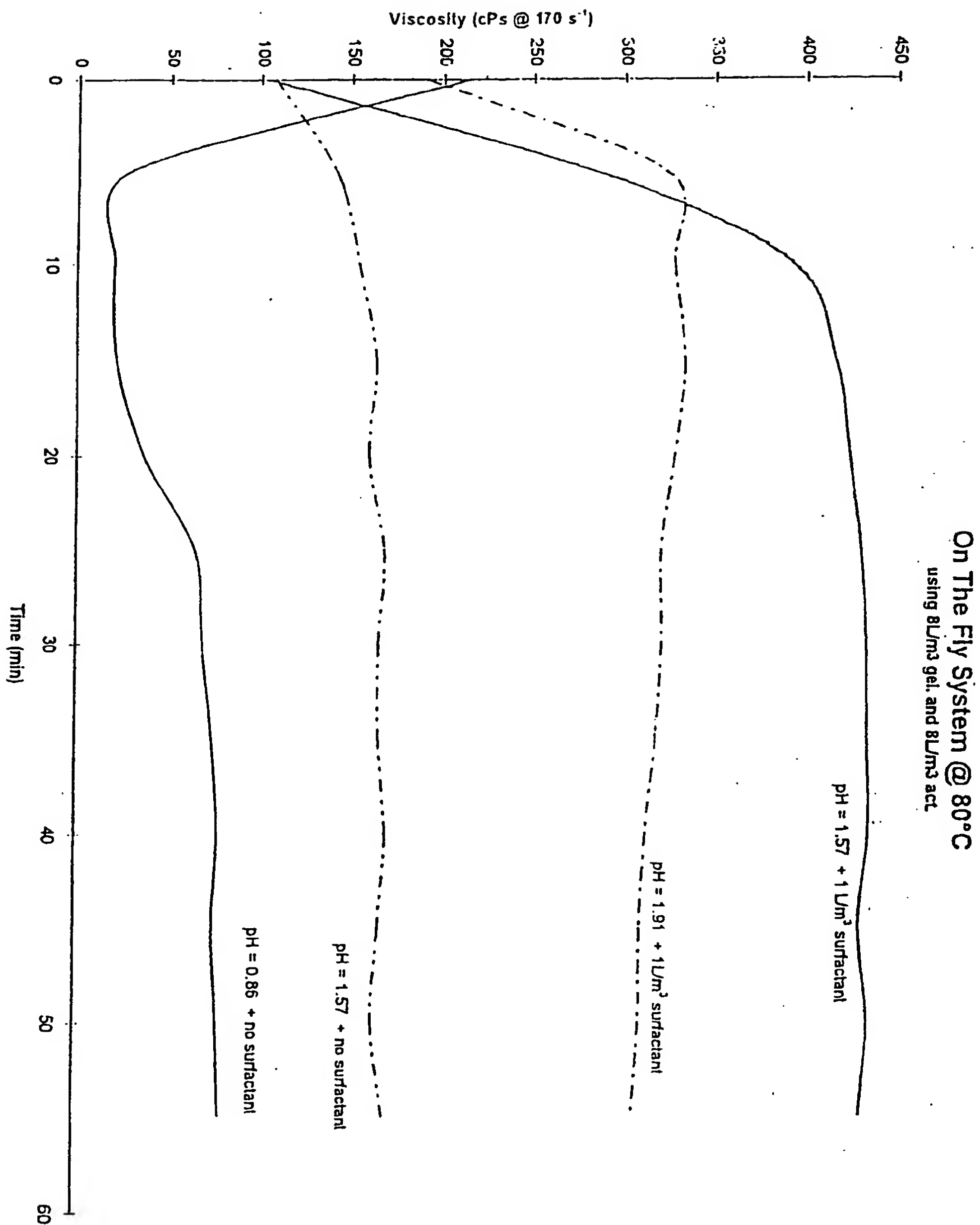


FIGURE 1

(19)



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(11)

EP 0 919 695 A3

(12)

EUROPEAN PATENT APPLICATION

(88) Date of publication A3:
20.10.1999 Bulletin 1999/42

(51) Int. Cl.⁶: E21B 43/26, E21B 43/25

(43) Date of publication A2:
02.06.1999 Bulletin 1999/22

(21) Application number: 98122163.3

(22) Date of filing: 25.11.1998

(84) Designated Contracting States:

AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE

Designated Extension States:

AL LT LV MK RO SI

(30) Priority: 25.11.1997 CA 2222087

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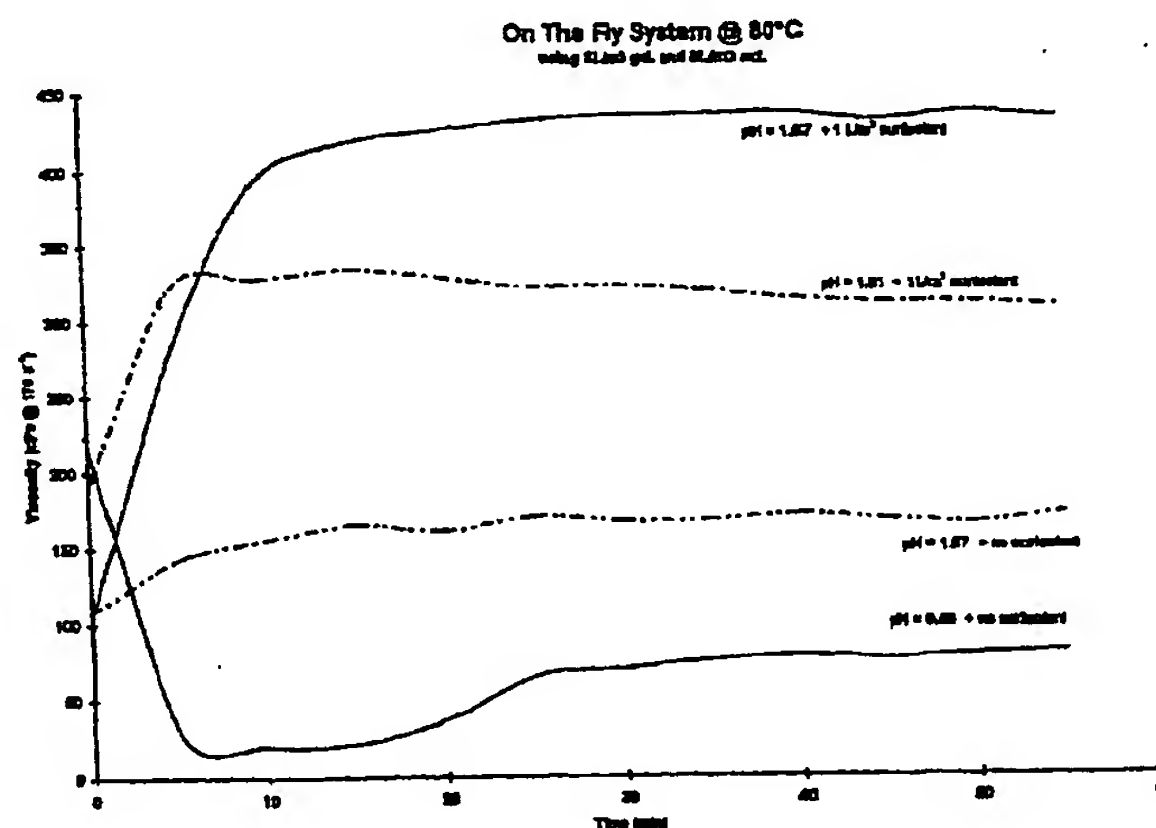


FIGURE 1

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EUROPEAN SEARCH REPORT

Application Number
EP 98 12 2163

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Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.6)
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X,P	DATABASE WPI Section Ch, Week 9912 Derwent Publications Ltd., London, GB; Class A18, AN 99-132863 XP002113207 & CA 2 232 776 A (NALCO/EXXON ENERGY CHEM LP), 14 October 1998 * abstract * ---	1	
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Place of search THE HAGUE		Date of completion of the search 24 August 1999	Examiner Boulon, A
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document</p>			

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**ANNEX TO THE EUROPEAN SEARCH REPORT
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EP 98 12 2163

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1. The first part of the document
describes the general situation
of the country.